

| <b>Negative Polyatomic Ions:</b>  |              |                           |                                  |
|---|--------------|---------------------------|----------------------------------|
| <b>Formula</b>  | <b>Name</b>  | <b>Formula</b>            | <b>Name</b>                      |
| $\text{NO}_3^-$   | nitrate      | $\text{ClO}_4^-$          | perchlorate                      |
| $\text{NO}_2^-$   | nitrite      | $\text{ClO}_3^-$          | chlorate                         |
| $\text{CrO}_4^{2-}$   | chromate     | $\text{ClO}_2^-$          | chlorite                         |
| $\text{Cr}_2\text{O}_7^{2-}$  | dichromate   | $\text{ClO}^-$            | hypochlorite                     |
| $\text{CN}^-$   | cyanide      | $\text{IO}_4^-$           | periodate                        |
| $\text{MnO}_4^-$  | permanganate | $\text{IO}_3^-$           | iodate                           |
| $\text{OH}^-$   | hydroxide    | $\text{IO}^-$             | hypoiodite                       |
| $\text{O}_2^{2-}$   | peroxide     | $\text{BrO}_3^-$          | bromate                          |
| $\text{NH}_2^-$   | amide        | $\text{BrO}^-$            | hypobromite                      |
| $\text{CO}_3^{2-}$  | carbonate    | $\text{HCO}_3^-$          | hydrogen carbonate (bicarbonate) |
| $\text{SO}_4^{2-}$  | sulfate      | $\text{HSO}_4^-$          | hydrogen sulfate (bisulfate)     |
| $\text{SO}_3^{2-}$  | sulfite      | $\text{HSO}_3^-$          | hydrogen sulfite (bisulfite)     |
| $\text{C}_2\text{O}_4^{2-}$   | oxalate      | $\text{HC}_2\text{O}_4^-$ | hydrogen oxalate (binoxalate)    |
| $\text{PO}_4^{3-}$  | phosphate    | $\text{HPO}_4^{2-}$       | hydrogen phosphate               |
| $\text{PO}_3^{3-}$  | phosphite    | $\text{H}_2\text{PO}_4^-$ | dihydrogen phosphate             |
| $\text{S}_2\text{O}_3^{2-}$   | thiosulfate  | $\text{HS}^-$             | hydrogen sulfide                 |
| $\text{AsO}_4^{3-}$   | arsenate     | $\text{BO}_3^{3-}$        | borate                           |
| $\text{SeO}_4^{2-}$   | selenate     | $\text{N}_3^-$            | azide                            |
| $\text{SiO}_3^{2-}$   | silicate     | $\text{CNO}^-$            | cyanate                          |
| $\text{S}_2\text{O}_8^{2-}$   | persulfate   | $\text{CNS}^-$            | thiocyanate                      |
| $\text{C}_2\text{H}_3\text{O}_2^-$ or $\text{CH}_3\text{COO}^-$             | acetate      |                           |                                  |
| <b>There are only three positive polyatomic ions that you need to know:</b> |              |                           |                                  |
| $\text{NH}_4^+$   | ammonium     | $\text{H}_3\text{O}^+$    | hydronium                        |
| $\text{Hg}_2^{2+}$  | mercury(I)   |                           |                                  |